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$-1R = 8.31 \text{ J K mol}^{-1}$ (from IB Data booklet) Use $PV = nRT$ (from IB Data booklet) $V = nRT/P = 1.50 \times 8.31 \times 298.15 / 1.08 \times 1.01 \times 10^5 = 30.034.1 \text{ m}^3 = 34.1 \text{ dm}^3$ 3[4] Award [4] for correct final answer. Award [3] (max) for 0.0341 dm³ or 22.7 dm³. 3Award [3] (max) for 34.4 dm³. 3Award [2] (max) for 22.9 dm³. 3Award [2] (max) for 0.0227 dm³.

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Catrin Brown teaches IB Chemistry at Lester B. Pearson United World College of the Pacific in Canada. She has been actively involved in the IB world since 1992 with extensive experience as an examiner, senior moderator, faculty member on the Online Curriculum Centre and a site visitor and consultant for schools.

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